

Suggest and explain the precautions necessary when potassium reacts with water.

Use the periodic table to identify and name a liquid in group 7

What do these 2 symbols mean?



How can we tell how many electrons an element has, using the periodic table?

What is the 'family name' for the elements in group 1?

What is the 'family name' for the elements in group 7?

Write balanced symbol equations for these reactions of alkali metals:

- a) Sodium with water
- b) Potassium with water
- c) Lithium with chlorine
- d) Sodium with chlorine

Use the periodic table to identify and name a solid element similar to chlorine

How many times heavier is

- a) a magnesium atom than a carbon atom?
- b) a sulphur atom than a helium atom?



Describe and explain the trend in reactivity in Group 1 as you go down the group

Describe and explain the trend in reactivity in Group 7 as you go down the group

C4 Chemical Patterns

Add notes and diagrams to make it useful

Complete this table:

Particle	Mass	Charge
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Electron

Proton

Neutron

Explain how spectroscopy is used to identify elements in chemical mixtures

What is the boiling point of water?

What are the formulae of these ionic compounds?

Magnesium oxide

Calcium chloride

Aluminium oxide

What are the 2 forms (species) that sodium can exist in?

In what ways are a fluoride ion, a neon atom and a sodium ion the same? How do they differ?

In aluminium oxide is made up of Al^{3+} ions and O^{2-} ions, what must its formula be?

Draw a diagram to show the electrons in shells for beryllium

Draw a diagram to show the electrons in shells for magnesium

When a metal atom loses an electron, what charge does it get?

When a non-metal atom gains an electron, what charge does it get?

Ionic salts can conduct electricity when either _____ or _____ but not when solid.

Which group of the periodic table do beryllium, magnesium and calcium belong to?

Draw the electron arrangements of chlorine (2.8.7)

Why do solid compounds made of ions not conduct electricity?

Match up names, charges and ions attracted.

ANODE	negative electrode	attracts + ions
CATHODE	positive electrode	attracts - ions

In what ways are a fluoride ion, a neon atom and a sodium ion the same? How do they differ?